Isotope → different # of n^0

p^+ → proton

n^0 → neutron

e^- → electron

Boron-11

atomic name - mass number

$= p^+ + n^0$

B $p^+ \rightarrow 5$

atomic # → 5

$n^0 \rightarrow 6$

mass # = $p^+ + n^0$

n^0 = mass # − p^+

Carbon-12

C 6 $p^+ = 6$ $n^0 = 6$ $e^- = 6$

neutral

Atomic Structure Practice:

Boron-11 B 5 5

$e^- = 2 + (-5) = -3$

$+2 = 5 + e^-$

Charge = $p^+ + e^-$

$+ \quad -$

$\emptyset = +5 + (-5)$

(neutral atom)

mass number = $p^+ + n^0$

n^0 = mass number $- p^+$